Electrocatalytic Reaction Induced Colloidal Accumulation: The Role of Dielectrophoresis

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ABSTRACT: A surface-driven flow is generated during the electrocatalytic reaction of a platinum–gold bielectrode within hydrogen peroxide. This flow can be experimentally visualized and quantified using micrometer-sized particles that are transported by a flow field. Tracer particles, which possess an inherent surface charge, also interact with the induced electric field and exhibit a collective behavior at the surface of the electrodes where they accumulate. The underlying mechanism for the accumulation dynamics demonstrated by these catalytic pump systems has so far been lacking. In this work, the accumulation dynamics and kinetics were experimentally investigated. With use of numerical simulations, we demonstrate that the self-driven particle accumulation is controlled by a positive dielectrophoretic force, mediated by the reaction-induced electric and flow field. These results contribute to the fundamental knowledge on immobilized bimetallic systems.

INTRODUCTION

Several biological systems exist in nature that convert chemical energy to execute high-precision translational and rotational autonomous motion. This has inspired a plethora of research and developmental activities on artificial micromachines, aimed at understanding and replicating the complex functionalities presented by these biological motors.1−4 A good example is the autonomous locomotion of a catalytic nanorod that functions primarily on the conversion of chemical energy harvested from its surrounding fluid media (see Figure 1a). The Pt–Au bielectrode catalyzes the decomposition of hydrogen peroxide via surface reactions, leading to the establishment of concentration and electrical potential gradients that propel the nanorod.

Previously, the locomotion of catalytic bimetallic motors was attributed to mechanisms based on interfacial tension5 and oxygen bubble recoil.6 Later, a consensus was reached on self-electrophoresis as the main driving force, which is controlled by localized concentration and electrical gradients and fueled by the chemical conversion of the surrounding electrolyte. The phoretic movement of the bimetallic motor is powered by a self-induced electric field in conjunction with the proton gradient that is generated during the electrochemical reaction, hence the name self-electrophoresis.

Paxton et al.7 initially demonstrated self-electrophoresis using striped bimetallic motors, which catalyzes the decomposition of hydrogen peroxide through surface reactions and creates electric field \( \mathbf{E} \) that propels the movements of the bimetallic motors (Figure 1a). The resulting electric field can induce particle migration whose velocity \( \mathbf{u}_{\text{ep}} \) can be expressed by the Helmholtz–Smoluchowski equation \( \mathbf{u}_{\text{ep}} = \mu_e \mathbf{E} \), where \( \mu_e \) is the electrophoretic mobility of the particle. The self-electrophoretic mechanism requires a net charged layer surrounding the rod, which is generated by the electrochemical reactions that generate the asymmetric distribution of charged intermediates. The transport of the ionic species is governed by diffusion, migration, and convective fluxes. The charge density combines with a self-induced electric field to create a body force that powers the rod movement.8 These fields, which are coupled, can be solved numerically.9 Moran and Posner10 developed a comprehensive self-electrophoretic numerical framework for a bimetallic motor by solving the coupled Poisson–Nernst–Planck and Stokes equations, where the surface reaction and fluxes can be accounted for using the simplified Frumkin corrected Butler–Volmer equations. The velocity of the bimetallic motor showed dependence on several parameters including the concentration of the fuel and conductivity of the medium.11

Kline et al.12 inspired by Paxton’s Au–Pt bimetallic motors,7 patterned an Au–Ag bielectrode on a substrate and contacted this with a hydrogen peroxide solution. This configuration immobilizes the bimetallic motor and restricts its motility...
during the electrochemical reaction with hydrogen peroxide. On the basis of the principle of Galilean relativity, a fluid flow was generated within the immediate surroundings of the bielectrode system. The immobilized bimetallic motor is also termed a micropump because of its fluid-pumping capability (see Figure 1b).

A catalytic micropump is a form of chemically powered pumps that exhibits diffusio-electroosmosis phenomena. Other types of chemically powered pumps can be activated through other mechanisms such as enzymatic reactions, photochemical reactions, and diffusioosmosis. A broad and in-depth review of chemically powered micropumps can be seen in Esplandiu et al. and Zhou et al.

The fluid flow is driven by an electric body force generated by an induced electric field acting on a net charge distribution, which arises from a proton concentration gradient that is established during the electrochemical reaction involving the decomposition of hydrogen peroxide. The reaction proceeds via oxidation and reduction pathways where oxidation occurs at the anode (platinum), leading to the generation of protons and oxygen, and reduction occurs at the cathode (gold), involving the consumption of protons and production of water. Early catalytic micropump designs were fueled by the decomposition of hydrogen peroxide. However, its low-conversion efficiency limits its potential for practical applications, which have necessitated the search for alternative electrolytes. Hydrazine derivatives consisting of N₂H₄ and N₂Me₂H₂ have been studied in conjunction with Au−Pd bielectrodes. More recently, halogen-based systems have been proven to be efficient and superior in performance compared to other forms of fuel-based systems. The chemomechanical actuation for these systems is similar to H₂O₂, where the decomposition of the fuel creates a gradient of charged intermediates that drive the fluid motion.

The induced fluid flow pattern is typically driven from the anode to the cathode and recirculated to the upper region within a confined space because of continuity. The fluid velocity u is impacted by the concentration and bulk pH of the electrolyte. The direction of the flow can be reversed from the cathode to the anode. This can be reaction-driven under lower anodic reactive regimes or by surface modification of the electrodes, which vary the zeta potential and impact the flow direction according to the Helmholtz−Smoluchowski formula $\mathbf{u} = -\frac{\zeta_m}{\eta} \mathbf{E}$, where $\zeta_m$ is the electrode’s zeta potential.

The flow field is commonly visualized with tracer particles that passively follow the fluid flow. An example of commonly used tracer particles include polystyrene beads, which can be functionalized with amine or carboxylate to tune the zeta potential of the particle. Such particles have been used to quantify the fluid velocity magnitude. The observed particle velocity magnitude $\mathbf{u}_p$ consists of the fluid flow component $\mathbf{u}$ combined with the electrophoretic velocity of the particle $\mathbf{u}_p$. 

![Figure 1](https://pubs.acs.org/doi/10.1021/acs.langmuir.1c01938)
The electrophoretic velocity is generated as a result of the interaction between the surface charge of the particles and the induced electric field. Decoupling these two velocity components will yield the actual magnitude of catalytically induced fluid flow. This can be achieved by using the two-particle correlation that involves the use of two tracers having different electrophoretic mobilities under the exact experimental conditions.25,29

In addition to particle transport by advection and electrophoresis, dynamic interactions and patterns have been observed for the tracer particles with electrocatalytic micro-pumps5,19,25 and micromotors.18 Wang and his team19,30 reported on the attachment of particles on bimetallic (Au−Pt) and trisegmented nanorods (Au−Ru−Au). The collective dynamics between these two classes of objects was attributed to an electrostatic force that is generated by the electric field.31 As soon as a moving nanorod approaches the charged particles, the particles are attracted toward the rod, leading to aggregation and formation of raft assemblies on the nanorod. Their modeling approach considered the nanorod as being uncharged, whereas the shear plane potential was made to float.

The dynamic interactions between an immobilized bimetallic system and colloidal particles occur irrespective of the combination of metals (Au−Pt, Au−Pd, Au−Ag, and Au−Cu) and fuel (H₂O₂, N₂H₄, and HCl) used in powering the system.5,19,32−34 In most cases, particles are observed to accumulate at the junction of the connected electrodes. The consensus for this behavior is that the particle attachment is primarily driven by an electrophoretic force. The induced flow transports the particles toward the surface of the electrodes where they become trapped and accumulate under the influence of the self-generated electric field.

The directed transport of colloidal particles has been demonstrated for applications involving cargo transport35 and biological assays,36 biosensors,37 and optics.38 In the aforementioned applications, the collective transport and accumulation kinetics of the particles are controlled by an externally applied field. In the case of immobilized electrocatalytic systems where the particle collective dynamics is self-driven, it is crucial to fully understand the underlying mechanism to fully exploit this behavior for relevant applications.

This article investigates the role of the dielectrophoretic (DEP) force toward the accumulation of charged particles in a bimetallic catalytic system. We present quantitative and numerical analyses that pertain to the agglomeration dynamics of particles at bimetallic electrocatalytic junctions for different time scales and electrolyte concentrations. The results show that the dielectrophoretic force, mediated by the electric field in conjunction with the surface-induced flow, plays a dominant role in particle agglomeration, which is observed for immobilized electrocatalytic systems.

THEORY AND NUMERICAL MODELING

Dielectrophoresis is an electrokinetic phenomenon exploited for several lab-on-chip applications that range from sorting, separation, manipulation, and concentration of micro-particles, in addition to cells and viruses.39−42 When a dielectric particle is exposed to a nonuniform electric field, it becomes polarized, and as a result, the surface charges are reoriented, inducing a dipole.43 The interaction between a dipole and an electric field gradient creates a dielectrophoretic force, F_{DEP}.44 For the immobilized electrocatalytic system, the dielectrophoretic force is strongly dependent on the reaction-induced electric field, which can be directly influenced by the geometry of the electrodes.

To understand and predict the interaction between the dielectrophoretic force and the tracer particles, we developed 2D models including Pt−Au bielectrodes. The bimetallic and interdigitigated electrodes are immersed into a solution of hydrogen peroxide and tracer particles and confined within a chamber of depth h. Figure 2 presents a schematic diagram of an immobilized Pt−Au bielectrode that catalyzes the decomposition of hydrogen peroxide via redox reactions. Oxidation occurs at the Pt end, which generates protons, oxygen molecules, and electrons. The electrons are directly transferred to Au to complete the circuitry. During the reduction reaction at the Au electrode, protons are consumed and combined with hydrogen peroxide to create water. The flux of protons between the electrodes establishes a proton concentration gradient that generates an electric field E and combines with the net charge distribution to form a body force that drives the interfacial fluid flow. The flow, which was visualized with tracer particles, is driven from platinum toward gold under normal electrocatalytic steady-state conditions and recirculates within a closed system.

The charged particles that are close to the surface of the bielectrode are trapped and concentrated at the junction of the
The dielectrophoretic force acting on the particle is expressed as
\[ F_{\text{DEP}} = 2\pi r^2 k_{\text{CM}}(\nabla \mathbf{E}^2) \] (1)
where \( \varepsilon \) is the fluid’s permittivity, \( r \) is the radius of a particle, and \( k_{\text{CM}} \) is the Clausius–Mossotti (CM) factor that is defined in terms of the complex permittivity of the particle \( \varepsilon_p^* \) and the fluid medium \( \varepsilon_f^* \) respectively as
\[ k_{\text{CM}} = \frac{\varepsilon_p^* - \varepsilon_f^*}{\varepsilon_p^* + 2\varepsilon_f^*} \] (2)

Under the influence of a direct current (DC) field or at frequencies below 100 kHz, the CM factor can be approximated in terms of the electrical conductivity of the fluid medium \( \sigma_f \) and the particle \( \sigma_p \) given as
\[ k_{\text{CM}} = \frac{\sigma_p - \sigma_f}{\sigma_p + 2\sigma_f} \] (3)
If \( \sigma_p > \sigma_f \), the dielectric particle will be attracted toward the region of high electric field strength (positive dielectrophoresis), and if \( \sigma_p < \sigma_f \) the particles are drawn to a weak electric field region (negative dielectrophoresis). According to O’Konski, the conductivity of a particle is defined as
\[ \sigma_p = \sigma_b + \frac{K_s}{r} \] (4)
where \( K_s \) denotes the surface conductance, which is about 1 nS for latex particles, and \( \sigma_b \) represents the bulk conductivity of polystyrene, which has been proven to be \( \sigma_b \approx 0 \). The surface conductance with particle radius \( r = 0.5 \mu m \) results in an effective particle conductivity of 40 \( \mu \)S/cm. The conductivity of the fluid medium was measured in the range of 5–30 \( \mu \)S/cm, leading to positive dielectrophoresis.

**Governing Equations.** Here, we present the governing equations that resolve the ion species concentration field, the induced electric field, the resulting flow field, and the trajectories of the colloidal particles. Moran and Posner used a similar approach to model the locomotion of catalytic nanomotors.

Under the assumption of a dilute solution limit, the concentration of ion species (H\(^+\), OH\(^-\), Cl\(^-\)) is solved by the Nernst–Planck equation,
\[ \mathbf{u} \cdot \nabla c_i = \nabla \cdot (D_i \nabla c_i + z_i F \mu_i \nabla \phi(c_i)) \] (5)
where \( \mathbf{u} \) represents the fluid velocity, \( c_i \) is the molar concentration for the \( i \)-th species, having a diffusion constant \( D_i \) with valence \( z_i \) and \( \phi \) is the electrostatic potential. The potential distribution in conjunction with the space charge density \( \rho_e \) is described by the Poisson equation,
\[ -\nabla^2 \phi = \rho_e / \varepsilon \] (6)
where \( \rho_e = \sum_{i=1}^{N} z_i F c_i \). The gradient of the potential distribution gives rise to the electric field described as \( \mathbf{E} = -\nabla \phi \).

The Stokes and continuity equations describe the fluid flow for an incompressible Newtonian fluid operating under a lower Reynolds regime (\( Re \ll 1 \)),
\[ 0 = -\nabla p + \eta \nabla^2 \mathbf{u} - \rho_e \nabla \phi \] (7)
\[ \nabla \cdot \mathbf{u} = 0 \] (8)

The movement of the particles is described by Newton’s law of motion,
\[ \frac{d(m_p)\mathbf{u}}{dt} = \mathbf{F}_d + \mathbf{F}_{\text{ep}} + \mathbf{F}_{\text{DEP}} \] (9)
where \( m_p \) is the mass of the particle and \( \mathbf{F}_d \) is the drag force that acts on the particles of radius \( r \) moving with velocity \( \mathbf{u} \) through a peroxide solution with viscosity \( \eta_p \) given as
\[ \mathbf{F}_d = \frac{1}{\tau_p} m_p \mathbf{M} (\mathbf{u} - \mathbf{v}) \] (10)
\( \tau_p \) is the response time of the particle velocity in seconds and \( \mathbf{v} \) is the velocity of the particles (m/s). \( M \) is a correction for the wall effects given as
\[ M = \left( 1 + \frac{1}{1 - \frac{9}{16} \alpha + \frac{1}{8} \alpha^2 - \frac{45}{256} \alpha^3 - \frac{1}{16} \alpha^4} \right) (1 - P(n)) \] (11)
where \( P(n) \) defines the projection operator onto the wall normal \( \mathbf{n} \), \( \alpha = \frac{l_n}{l_s} \), and \( l_s \) is the distance from the center of the particle to the nearest wall.

The electrophoretic force \( \mathbf{F}_{\text{ep}} \) is defined as
\[ \mathbf{F}_{\text{ep}} = 6 \pi \eta_p \mathbf{u}_{\text{ep}} \] (12)
where \( \mathbf{u}_{\text{ep}} = \frac{z \varepsilon_e}{\varepsilon_k} \mathbf{E} \) is the electrophoretic velocity of the particles. The zeta potential of the particles is denoted by \( \zeta_p \), which was measured to be \( -45 \text{ mV} \).

The drag, electrophoretic, and dielectrophoretic forces combine to a net resultant force on the particles, resulting in its acceleration. The particles in this case are assumed to be spherical in shape. The gravitational force is neglected because the particles are assumed to be neutrally buoyant.

**Boundary Conditions.** A no-slip condition, \( \mathbf{u} = 0 \), is applied to the top and bottom boundaries. A symmetry boundary condition is applied on the left and right walls of the bimetallic system model. For the interdigitated electrode, we imposed a periodic boundary condition on the left and right walls to indicate a repeated pattern for Pt and Au for which we simulate a single period. For the potential boundary conditions, we solved explicitly the diffuse part of the electric double layer. The potential drop across the Stern layer is described by
\[ \Delta \phi_{\text{Stern}} = \zeta - \phi_{\text{electrode}} \] (13)
which scales with the ratio of the Stern layer length \( L_{\text{Stern}} \) to the Debye length \( L_D \) as well as the zeta potential of the electrode as
\[ \Delta \phi_{\text{Stern}} \approx \frac{L_{\text{Stern}}}{L_D} \zeta_{\text{electrode}} \] (14)
where \( L_D = \sqrt{eV_T/2Fc_0} \) is usually hundreds of nanometers, \( V_T = RT/F \approx 25 \text{ mV} \) is the thermal voltage, \( T \) is the temperature, \( F \) is the Faraday constant (\( F = 96485 \text{ C mol}^{-1} \)), \( R \) is the universal gas constant (8.314 J/K mol), and \( c_0 \) represents the bulk proton concentration. For an electrolyte having a low salt...
Numerical Method and Simulation. The 2D models were partitioned into subdomains to implement a user-controlled nonuniform mesh that consists of triangular elements. A uniform mesh was imposed at the tangential x-direction within the lower part of the model, using mapped meshing control. This results in refinements that resolve the electric double layer and other gradients at the surface of the electrodes.

The governing equations are solved using COMSOL Multiphysics V.5.5. At the beginning of the simulation, the steady Nernst–Planck and Poisson equations are computed without the flow field to generate the initial conditions for the system, which were used in solving the coupled governing equations sequentially until the solution converges. The results are implemented as initial values for a time-dependent solver used in resolving the particle trajectories, where time stepping is performed using a generalized alpha implicit method.

EXPERIMENTAL METHOD

The colloidal accumulation experiments were performed using platinum–gold bielectrodes that are immobilized on glass substrates. The bielectrodes were fabricated as follows: a positive resist was spin-coated on a glass wafer and then was exposed to UV light through a photomask and development. Both platinum and gold electrodes were sputtered on tantalum (5 nm), which was predeposited on the glass wafer (MESA+ nanolab cleanroom in-house equipment, “TCOathy”). Undesired metal residues were removed during the lift-off process by ultrasonification in acetone. Both Pt–Au interdigitated (IDE) electrodes and Pt–Au bimetallic electrodes were fabricated by this method (see Figure 3).

The interdigitated electrodes were bonded to a printed circuit board, thereby forming a configuration that allows us to control the electrochemical reaction. For the bimetallic electrodes, Pt and Au constitute a galvanic pair that maintains continuous contact with each other. Hydrogen peroxide (Sigma-Aldrich) solutions of known concentrations were prepared in Milli-Q water (ρ > 18 MΩ cm). The pH of the resulting solutions were adjusted accordingly using HCl in the range pH 5–6, and the conductivity was measured with a conductivity probe (WTW Cond 3110, Weilheim, Germany) to be in the range of 5–30 μS/cm. Particle dispersions were prepared by seeding the hydrogen peroxide (10 mL) solution with 2.5 wt % (1 μL) fluorescent microparticles (PS-FluoRed-Fi329 by Microparticles GmbH), 1 μm in diameter, and introduced into a confined hybrid chamber that contains the Pt–Au bimetallic electrode. The motion and accumulation of the particles are observed with an inverted optical microscope (Carl Zeiss Axio Observer Z1, 20X). Image frames were recorded with a CCD camera with a frame rate of 10 fps.

Image analysis was done using the open access software ImageJ. The image frames were digitized by converting to 8 bits. The contrast and brightness were adjusted, and segmentation was performed to isolate the particles by subtracting the image background. The images were transformed to a binary form by applying a standard maximum entropy threshold. Particle analysis was performed using the “analyze"
command, which determines the number of particles as well as the average size of the accumulated particles in each image frame.

RESULTS AND DISCUSSION

Experimental Results. The movement and collective dynamics of the particles are observed with optical microscopy. The tracer particles are transported near the surface of the electrodes (at a relative height around $y = 10 \mu m$) by the induced fluid flow from the anode (platinum) to the cathode (gold) (see Movie 1 Supporting Information). The particles close to the junction of the connected electrodes become trapped, and the population grows in clusters (see Movie 2 Supporting Information). Figure 4 shows the time series for the particle accumulation at the junction of a Pt–Au bimetallic electrode, which intensifies with time and slightly extends above the surface of the bi-electrode.

The particle accumulation behavior was related to the reaction mechanism involving the catalytic decomposition of hydrogen peroxide by repeating the experiment using interdigitated electrodes, where the reaction is monitored and controlled by an external current measurement. During the connected mode, the particles close to the surface of the electrodes are migrating toward the platinum electrode where they maintain their Brownian motion (see Figure 5a, Movie 3 Supporting Information). At longer time scales, the particles are observed to accumulate and grow into clusters (see Movie 4, Supporting Information). As soon as the electrodes are disconnected, the particles are released from their trapped state and slowly disperse again into the bulk region (Figure 4, Movie 5 Supporting Information). To understand what is happening, we refer to the electrochemical behavior of the Pt–Au interdigitated electrodes in hydrogen peroxide solution. The thus-generated electric field plays two main roles. First, it drives the fluid flow in conjunction with the space charge density in an electroosmotic fashion. Second, it interacts with the particles to create (dielectro)phoretic forces that affect the movement of the particles (Figure 2).

Now that we know that the particle accumulation dynamics is controlled by the reaction-induced electric field, we quantify the accumulation dynamics by estimating the surface coverage (the area occupied by the trapped particles vs the total surface area of the electrode) of the trapped particle monolayer formed during the first minutes of the accumulation process involving the Pt–Au bimetallic electrode. The quantification method is described in section S1 of the Supporting Information. The surface coverage is plotted as a function of time for different $\text{H}_2\text{O}_2$ concentrations (0.05, 0.108, and 0.163 M) and shown in Figure 6a. The clear trend depicts the
increase in the accumulation rate with an increase in the \( \text{H}_2\text{O}_2 \) concentration.

To elucidate the influence of \( \text{H}_2\text{O}_2 \) on the accumulation kinetics, we determined the agglomeration rate constant from the number of trapped particles \( n_p \) versus time, which is fitted with a linear regression. This was evaluated during the first 5 min of the accumulation process before the particles form multilayers. Figure 6b suggests that the accumulation rate increased rapidly at higher peroxide concentrations owing to the increase in field strength due to the impact of the peroxide concentration on the electrochemical reaction rate.

Particle Accumulation by DC Dielectrophoresis. A series of control tests was performed to ascertain the dominance of the dielectrophoretic force on the particles. This was done by varying the conductivity of the fluid that results in the variation of the CM factor and plotted against the particle agglomeration rate (see Figure S2.1, Supporting Information). As the conductivity of the suspending medium increases, the CM factor decreases and the particle accumulation rate driven by the positive DEP force reduces. This trend is also observed in the simulation results obtained by varying the conductivity of the fluid, where an increase in the conductivity of the fluid medium results in the reduction of trapped particles by the DEP force.

Our numerical simulations provide more insights on the particle accumulation behavior at the surface of the electrodes. We start by analyzing the reaction-induced electric field that drives the particle accumulation. Figure 7 shows the electric field lines across the Pt−Au bimetallic and interdigitated electrodes, combined with the reaction-induced potential distribution (color) that was evaluated for 0.163 M \( \text{H}_2\text{O}_2 \), pH 6.1. In both cases, the electric field is generated by the proton gradient that originates from the electrochemical reaction involving the decomposition of \( \text{H}_2\text{O}_2 \) and extends from the double-layer region just after the surface potentials of the electrodes (see the insets of Figures 7a,b) to the bulk region. The reaction-induced
potential acts to balance the reactive current from the electrodes through the variation of the proton concentration at the electrode’s surface.\textsuperscript{23}

We see in Figure 7 that the electric field originates from platinum and the field line extends toward gold. The apparent source of the electric field, which is rooted at platinum, is the region where the field contributions from the reaction-induced proton gradient and surface charges intersect.\textsuperscript{10} The direction of the field lines is dictated by the proton current across the bielectrode. Figure 8\textsuperscript{a} shows the tangential electric field profile along the length of the bimetallic electrode whose gradient is maximum at the junction where both electrodes maintain a continuous contact. The physical inhomogeneity at the midpoint increases the electric field strength that creates the DEP force. The electric field combined with the charge density drives the fluid flow that transports the particles near the surface of the electrode during recirculation (Figure 8\textsuperscript{b}), within the range of the positive dielectrophoretic force where
they become trapped and accumulate. The charged particles are capable of generating electroosmotic flow due to interactions between the induced electric field and the double layer of the surface of the particles, which act to entrain nearby particles on the electrode. But they were not computed in the model as they are hardly visible in this case (see Movie 1 and Movie 2 in the Supporting Information).

Different strategies have been previously adopted to increase the local field to concentrate particles within the high field region. A common method is to use insulating structures that generate spatial nonuniformity in the local electric field known as insulator-based dielectrophoresis (iDEP), which increases the intensity of the field that strongly affects the particles. In another instance, a scratch was introduced on the surface of an ITO electrode, which increased the local current density by a factor of 2 compared to that of unscratched electrodes, and promotes a higher field strength that traps colloidal particles within the scratched region.

In the case of the interdigitated electrode, the tangential electric field is maximum at the end of the platinum electrode just before the spacing between Pt and Au, and the resulting flow streamlines that are driven by the field gradient build up from this region, which transport particles close to the platinum end of the interdigitated electrode where they eventually accumulate (see Figure 3.1).

Figure 9a shows the average distribution of forces acting on the particles that are plotted from the bielectrode junction along the channel height. At several micrometers from the surface region, \( y > 1 \) \( \mu m \), the electrophoretic and drag forces dominate the DEP force. At \( y \leq 1 \mu m \), the DEP force is amplified by the high electric field strength and gradient and becomes more significant. At the surface region, the DEP force surpasses the other two forces by almost 3 orders of magnitude. As a result, the particles that are transported by the flow field are impacted by both the drag force and the electrophoretic force, and as soon as they are within the range of the positive DEP force, they are pulled toward the region of the maximum local electric gradient where they accumulate. The direction of these forces are highlighted by the vector plot in Figure 9b, and the magnitude of the DEP force is greater compared to that of the other two near the surface region of the bielectrode. The maximum of each force contribution was determined for different electrolyte conditions and plotted versus the concentration of \( \text{H}_2\text{O}_2 \) (Figure 9c). All forces were greater than the Brownian force acting on the particle, which was determined to be on the order of magnitude \( 10^{-11} \) N (see Supporting Information S4). We see that the positive DEP force exerts the most dominant force on the particles, which occurs at the bielectrode junction.

The accumulation statistics for the trapped particles at the junction of the Pt–Au bimetallic electrodes is determined numerically and compared to our experimental results for different \( \text{H}_2\text{O}_2 \) concentrations (see Figure 10). The accumulation proceeds at a very fast rate, making it challenging to determine the trapped particle statistics for the first few seconds experimentally, which results in deviations at \( t < 100 \). In the numerical results the number of trapped particles and the agglomeration rate is influenced by the number of initial particles, \( N \), used for the simulation (see Figure S2.2, Supporting Information S2); nonetheless, the numerical result matches well with the experimental data in Figure 10 when the initial number of particles is set to 1000. At longer time scales, the model is capable of highlighting localized particle trapping and accumulation exhibited by bimetallic catalytic micro-pumps.

**CONCLUSION**

In this article, we studied the dynamics of the reaction-driven colloidal trapping and accumulation by a Pt–Au bimetallic catalytic system. We showed that the accumulation of particles occurs under the influence of a positive dielectrophoretic force that is mediated by a self-induced electric field. The particle accumulation kinetics were studied experimentally and found to scale with the concentration of the fuel that drives the electrocatalytic reaction, which generates the induced electric field. Numerical modeling elucidates the role of the drag and electrophoretic force that directly influences the movement of the particles further away from the surface region, while they accumulate under the effect of a positive dielectrophoretic force at the surface. Our combined experimental and numerical approaches shed light on the underlying mechanism of the reaction-induced particle accumulation and advance the knowledge of the transport mechanism by immobilized electrocatalytic systems.

**ASSOCIATED CONTENT**

**Supporting Information**

The Supporting Information is available free of charge at https://pubs.acs.org/doi/10.1021/acs.langmuir.1c01938.

Particle agglomeration quantification method; DC dielectrophoresis control tests and results; simulation results showing field quantities (velocity, potential, and dielectrophoretic force) for the interdigitated electrodes; Brownian force acting on particles (PDF)

Movie 1: Transport of colloidal particles by the catalytically induced flow field from platinum to gold (MP4)

Movie 2: Time lapse for particle trapping and accumulation at the junction of the Pt–Au bimetallic electrode (MP4)

Movie 3: Accumulation of particles on the surface of the platinum electrode during the connected mode for the interdigitated electrodes (MP4)

Movie 4: Accumulation of particles into clusters at longer time periods (MP4)
Movie 5: Behavior of the particles under the disconnected mode of operation; the trapped particles break into singlelets and slowly disperse back into the bulk region (MP4)

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Notes
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